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704 CHAPTER 22 Identify the product that balances the following nuclear reaction: ${}_{84}^{212}\text{Po} \rightarrow {}_{2}^4\text{He} + 1$. The total mass number and atomic number must be equal on both sides of the equation. ${}_{84}^{212}\text{Po} \rightarrow {}_{2}^4\text{He} +$ mass number: 212

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~~208~~ atomic number: 84 ~~2~~ =82 2. The nuclide has a mass number of 208 and an atomic number of 82, $^{208}_{82}\text{Pb}$.
3.

~~CHAPTER 22 Nuclear Chemistry~~

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Veritas/ Chemistry/ Chapter 4 Answer Key 3 Practice Problems 13. Use dimensional analysis and the given densities to make the following conversions. a. 14.8 g of boron (B) to cubic centimeters of boron. The density of boron is 2.34 g/ cm³. $(14.8 \text{ g}) \left(\frac{1 \text{ cm}^3}{2.34 \text{ g}} \right) = 6.32 \text{ cm}^3$ b. 2.8 L of argon (Ar) to grams of argon.

~~Chapter 4 Problem Solving in Chemistry Answer Key~~

1) No, it is not a proper answer; you do not know whether the professor meant homework problem number 20 or 20 homework problems. Twenty is a quantity and does not make sense without a unit. 3) Middle Zeros: Zeros that appear between other nonzero digits are ALWAYS significant.

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