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- Write the equilibrium expression. Look up the value for K_a (in a table).
- Write down the initial and equilibrium concentrations for

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everything except pure water. •We usually assume that the equilibrium concentration of H^+ is x . •Substitute into the equilibrium constant expression and solve.

•Remember to convert x to pH if necessary.

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system in equilibrium
. Find the reaction .
quotient, Q , (K_{eq}) for
a reaction equation .
Solve for Q given .
initial concentrations
or pressures. Using .
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determined by the

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"dissociation"

equilibrium. $\frac{3}{4}$ Strong

acids/strong bases 1.

dissociates

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completely in water

2. have very large
dissociation or K
values $\frac{3}{4}$ Weak

acids/weak bases 1.

dissociate only to a
slight extent in water
2.

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