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acids- compounds that produce an increase in  $[H^+]$  when dissolved in water. bases- compounds that produce an increase in  $[OH^-]$  when dissolved in water Lewis Definitions:  
acids- electron pair acceptors.

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aqueous solution between hydrochloric acid, a strong acid that completely dissociates to produce  $H_3O^+$ , and sodium hydroxide, a strong base that completely dissociates to produce  $OH^-$ . The formula equation for this reaction is written as follows.

$$HCl(aq) + NaOH(aq) \rightarrow NaCl(aq) + H_2O(l)$$

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Acid and Base Worksheet - Answers.

1) Using your knowledge of the Brønsted-Lowry theory of acids and bases, write equations for the following

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acid-base reactions and indicate each conjugate acid-base pair: a)  $\text{HNO}_3 + \text{OH}^-$  ( $\text{H}_2\text{O} + \text{NO}_3^-$ ).  $\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other. b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O}$  ( $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$ )

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and bases change the color of  
compound indicators. ... -Lewis acid-  
base reaction is the formation of one  
or more covalent bonds between an  
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pair acceptor, although the 3 acid-base reactions differ, many compounds can be categorized as acids ...

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Choice Questions 1) \_\_\_\_\_ is found in carbonated beverages due to the reaction of carbon dioxide with water.

A)  $\text{CH}_3\text{COOH}$ ; B)  $\text{H}_2\text{CO}_3$ ; C)  $\text{HCOOH}$ ; D)  $\text{C}_6\text{H}_5\text{COOH}$ ; E)  $\text{CH}_3\text{CH}_2\text{COOH}$ ; Answer: B. Diff: 1

Page Ref: 15.2 2) Which of the following species is amphoteric? A)

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CO<sub>3</sub><sup>2-</sup>; B) HF; C) NH<sub>4</sub><sup>+</sup>; D) HPO<sub>4</sub><sup>2-</sup>

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completely in solvent. a base that ionizes completely in a solvent. an acid that releases few hydrogen ions in aqueous solution. a base that releases few hydroxide ions in aqueous solution.

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The reaction is as follows:  $ZnO(aq)$

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$+ 2 \text{HCl} (aq) \rightarrow \text{ZnCl}_2 (aq) + \text{H}_2\text{O} (l)$  Zinc oxide reacts with sodium hydroxide to produce zincate sodium and water. This reaction takes place at a temperature of  $500-600^\circ\text{C}$ . It is an exothermic reaction.  $\text{ZnO} + 2 \text{NaOH} \rightarrow \text{Na}_2\text{ZnO}_2 + \text{H}_2\text{O}$ . g. Limestone is calcium carbonate.

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