

Chapter 18 Reaction Rates And Equilibrium Answer Key

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Chapter 18 - Reaction Rates and Equilibrium

a state of balance in which the rates of the forward and reverse reactions are equal; no net change in the amount of reactants and products occurs in the chemical system (18.2)

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The rate of a chemical reaction is expressed as the amount of reactant changing per unit time. Work Step by Step The rate of a chemical reaction is expressed as the amount of reactant changing per unit time.

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CHAPTER 18. Reaction Rates and Equilibrium. 18.1 Rates of Reaction. •The time needed to complete a

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chemical reaction can vary tremendously. •Rates measure changes that occur within intervals of time. •In chemistry, rates or speeds are expressed in chemical terms. Atoms, ions, and molecules must collide in order to react. Atoms, ions, and molecules can form a chemical bond when they collide, as long as the particles have enough kinetic energy and have the proper orientation.

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chapter 18--- reaction rates and equilibrium. STUDY. PLAY. rate. describes the change over an interval of time. collision theory. atoms, ions, and molecules can react to form products when they collide, provided that the particles have enough kinetic energy. activation energy.

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