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III Chapter 5 Electrons  
in Atoms Pt II Electron  
Configuration — Basic

~~introduction~~ The  
Electron: Crash Course  
Chemistry #5

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Quantum Numbers,  
Atomic Orbitals, and  
Electron Configurations  
Valence Electrons and  
the Periodic Table Intro  
to Ch. 5: Electrons in  
Atoms ~~Ch 5 Sec 1~~

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~~Atoms in Electrons~~

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Bohr Model of the  
Hydrogen Atom,  
Electron Transitions,  
Atomic Energy Levels,  
Lyman \u0026amp; Balmer  
Series ~~Atoms | What are  
They? What are  
Protons, Neutrons and  
Electrons? What Is An  
Atom?~~

---

The Photoelectric Effect  
~~Atoms and Molecules~~  
~~Class 9 Tutorial~~ How to

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write electron configurations and what they are How Small Is An Atom? Spoiler: Very Small. How to find the number of protons, neutrons, and electrons from the periodic table

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Section 1: Organizing the Elements ~~Energy from Wavelength: Electromagnetic Radiation Calculation~~

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IB Chemistry Topic 2

Atomic structure 12.1

Electrons in atoms HL

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~~Section 2: Electron~~

~~Arrangements in Atoms~~

~~Quantum Numbers~~

~~The Easy Way!~~

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Atomic Structure And

Electrons - Structure Of

An Atom - What Are

Atoms - Neutrons

Protons Electrons

---

Pearson Chapter 5:

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Section 1: Revisiting the  
Atomic Model Ch 5

Electrons in Atoms pt 1

Chapter 9 - Electrons in  
atoms and the Periodic  
Table Chapter 5

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Chemistry by

Ms.Basima Chapter 5

Electrons In Atoms

138 Chapter 5 □

Electrons in Atoms

Although the speed of  
all electromagnetic



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waves in a vacuum is the same, waves can have different wavelengths and frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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SmileyKylie0923. Key

Concepts: Terms in this  
set (57) Dalton. The

atom is a tiny,

indestructible particle

with no internal

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structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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5.1 Wave-Particle  
Duality/Electromagnetic  
Spectrum/Relationship  
of

Wavelength, Frequency  
and Speed of light 5.2

Bohr's Model of the  
Atom/Quantum

Mechanical Model of  
the Atom 5.3 Electron  
Arrangement & Valence  
Electrons.

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Chapter 5: Electrons in  
Atoms Models of the  
Atom Rutherford used  
existing ideas about the  
atom and proposed an  
atomic model in which  
the electrons move  
around the nucleus, like  
the planets move around  
the sun. Rutherford's  
model fails to explain

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why objects change  
color when heated.

Atoms Study

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Section 5.2 □ Electron

Arrangement in Atoms

The electron

configuration of an atom

is the arrangement of

the electrons. There are

3 rules that govern the

electron configuration:

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Aufbau's principle,  
Pauli Exclusion  
principle, and Hund's  
rule.

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Chapter 5 □ Electrons in  
Atoms □ Chemistry  
Charles Page High  
School Stephen L.  
Cotton \* \* \* \* \*

\* \* \* \* \* The  
electromagnetic  
spectrum consists of  
radiation over a broad  
band of wavelengths.  
The visible light portion  
is very small. It is in the



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10-7m wavelength  
range and  $10^{15}$  Hz (s<sup>-1</sup>)  
frequency range.

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Electrons In

Unit 4: Electrons in  
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You'll Remember ...

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Quantum Mechanical  
Model. model of the  
atom we believe today  
that involves the  
probability of finding an  
electron in a certain  
position. What is the  
maximum number of f

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orbitals in any single energy level in an atom ? 7. Electrons in the same orbital.

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in Atoms CHAPTER 5

What You'll Learn You

will compare the wave

and particle models of

light. You will describe

how the frequency of

light emitted by an atom

is a unique characteristic

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of that atom. You will compare and contrast the Bohr and quantum mechanical

## Chapter 5 Electrons In Atoms

138 Chapter 5 Electrons  
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Configurations for  
Elements in Period  
Three Table 5-4 Figure  
5-19. This sublevel  
diagram shows the order

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in which the orbitals are usually filled. The proper sequence for the first seven orbitals is 1s, 2s, 2p, 3s, 3p, 4s, and 3d. Chapter 5 Electrons in Atoms Flashcards | Quizlet

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Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun.

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Atoms ...

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Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Answers

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What You'll Learn You  
will compare the wave  
and particle models of  
light. You will describe  
how the frequency of

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Light emitted by an atom is a unique characteristic of that atom. You will compare and contrast the Bohr and quantum mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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How many electrons can each p orbital hold?

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times. 0. 10th - 11th

grade .

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Chapter 5 Electrons in

Atoms 2. Light and

Quantized Energy (5.1)

- The study of

light led to the

development of the

quantum mechanical

model.

- Light

is a kind of

electromagnetic

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## Chapter 5

radiation EM).

- All move at  $3.00 \times 10^8$  m/s
- (c) Speed of light.
- 3.

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