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Answers . . .

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Covalent Bonds

1. A covalent bond forms when two or more valence electrons are attracted by the positively charged nuclei of two atoms and

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thus are shared between both atoms. 2. The H₂ molecule is stable because each hydrogen atom now has a shared pair of electrons and has achieved a stable noble gas configuration. 3.

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used to measure

the number of

atoms,

molecules, or

(in the case of

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formula units in
a given mass of
a substance.

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Complete each
statement below
by choosing a
term from the
following list.

... Concept

Review: Covalent

Bonds 1. A

covalent bond

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or more valence
electrons are
attracted Holt
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4. stances. 2.

Answers will vary but may include iron rusting, milk. .

.

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number of atoms,
molecules, or
(in the case of
ionic compounds)
formula units in
a given mass of
a substance. The
mole is defined
as the amount of
substance that
contains the
number of carbon
atoms in exactly
12 g of

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Bonds-12,
Avogadro's
number (6.022×10^{23}) of atoms
of carbon-12.

3.2: The Mole
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Compounds Name

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continued Use

the system of

prefixes and the

rule for

suffixes to name

the following

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PbCl₂ 16.

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term from the

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Compounds Ionic

Compounds. The

strongest force

between any two

particles is the

ionic bond, in

which two ions

. . .

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Identify the
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help. Find the
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electric
current, the
substance must
possess free-
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particles. These
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delocalized
electrons, such
as those found
in substances
that form
metallic bonds.

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1. Use the
concept of
potential energy
to describe how

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a covalent bond forms between two atoms. As the atoms

Answers

involved in the formation of a covalent bond approach each other, the electron-proton attraction is stronger than the electron-electron and

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repulsions. The
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and their
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where the

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electrons are shared equally is called a nonpolar bond (eg H-H) and an unequal sharing of the pair of bonding electrons results in a polar bond. The unequal sharing of electrons is due to the

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ability of an
atom to attract
electrons
towards itself
which is known
as Electronegati
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chemical bonding theory that is based on the assumption that an atom, radical, or molecule that has gained or lost one or more electrons is an ion that has a positive charge due to losing electrons and an ion that has a negative

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atoms tend to have either empty valence shells or full valence shells of eight electrons. Ion. an atom, radical, or molecule that has gained or lost one or more electrons and has a negative

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