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Why? Many of the items you use on a daily basis rely on batteries for power. ... Key Questions 1. For neutral compounds, the sum of all oxidation numbers must equal _____ .

... If yes, use specific examples from the

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model to support your answer. 10. In the space under each reaction in the model, write ...

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2 POGIL™ Activities for High School Chemistry

3. In the space under each reaction in Model

1, write the oxidation number for every atom.

Divide the work among your group members. An

example is shown here: $4\text{Fe}(s) + 3\text{O}_2(g) \rightarrow 2\text{Fe}_2\text{O}_3(s)$?

0 0 +3 -2 0 +3 -2 -2 4.

Identify any elements that changed oxidation

number in the reactions in Model 1.

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LEO = Loss of Electrons is Oxidation 1 1.
Consider the incomplete half-reactions below.
a. Use oxidation numbers to identify the

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reactions below as oxidation or reduction. b. Place the correct number of electrons on the appropriate side of the reaction to complete the equation. PEOVCTI&..J 6 x STOP CF. e⁻ 12. Consider Reaction A in Model 2.

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