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Many companies, including pharmaceutical companies, keep track of the concentration of a solution by measuring its molarity – a ratio of the number of solute particles to the volume of the solution. In this activity you will learn about molarity and how to represent concentration quantitatively. Model 1 – Lemonade Mixtures *

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Pogil Extension Questions Molarity Answers. The most common measure of concentration used by chemists is molarity (M). Molarity is defined as the number of moles of solute (mol) divided by the total volume (V) of the solution in liters (L). Molarity = moles of solute per liter of solution (M = mol / L).

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of the new solution? (Answer: 9-801 HNO3) = ML (0.8M -N 3-30.00 ml, of water is added to 20.00 ml, of 0.0500 M sodium carbonate. What is the molarity of sodium carbonate in the new solution? (Answer: 0.0200 M) (20.00 m C (350. 00 casF) 4. What volume of 0.100 M NaCl is needed to make 5.00 ml, of 0.200 M solution? (Answer: 0.01 (0.100 M) V)

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