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integral over  $\Omega$  is:  $\sigma^2$ , so the standard deviations of molecular speed are  $\sigma^2 = (k_B T/m)^{1/2} = \sigma^2 = \sigma^2$ , where the second two equalities follow from repeating this calculation for the second and third directions. b) From (1.27),  $nV = p/k_B T = \rho / m = \rho / a$ , and  $\rho = \rho / a$ .

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