Study Guide The Mole Measuring Matter Answers

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The mole is the most common unit used to express the quantity of a chemical substance. For all solids, liquids, and gases, you can convert mass to moles (or moles to mass). For gases, you should memorize the following conversion of volume to moles (or moles to volume): at 0°C and 1 atmosphere pressure (known as standard temperature and pressure or STP), one mole of any gas occupies approximately 22.4 liters.

The Mole Unit - CliffsNotes Study Guides
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A mole of a particular substance is equal to the number of atoms in exactly 12 g of the carbon-12 isotope. Experiments established that number to be 6.022142 * 10 23 particles. Avogadro's number = NA = 6.022 * 10 23 items/mole. In other words, one mole equals 6.022 * 10 23 items. The units of the mole are modified depending on the units needed in a calculation.

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6.1.81 1024. 7.16.05 g/mol. 8.1.20 1024. Section 10.4 Empirical and Molecular Formulas. 1. The percent composition of a compound is the percent by mass of each of the elements in a compound. 2. Divide the mass of each element in the sample by the mass of the sample. Then multiply each quotient by 100.

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Read Book Study Guide The Mole Measuring Matter Answers and is the SI unit for measuring the amount of a substance. The number of representative particles in a mole, 6.02 10 23 , is called Avogadro's number . 10.1 The Mole: A Measurement of Matter 10 Chapter 10 Chemistry Study Guide. When you are measuring at STP the volume of any gas will ...

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